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Name: KEY
Date: _____

Solve the Equations for y

$$1) \begin{array}{r} -3x + y = -6 \\ +3x \quad +3x \\ \hline y = 3x - 6 \end{array}$$

$$6) \begin{array}{r} -2x - y = 4 \\ +2x \quad +2x \\ \hline -y = 2x + 4 \\ \hline y = -2x - 4 \end{array}$$

$$2) \begin{array}{r} 3x - 2y = 6 \\ -3x \quad -3x \\ \hline -2y = -3x + 6 \\ \hline y = \frac{3}{2}x - 3 \end{array}$$

$$7) \begin{array}{r} 12 + y = 4x \\ -12 \quad -12 \\ \hline y = 4x - 12 \end{array}$$

$$3) \begin{array}{r} 9x + 3y = -12 \\ -9x \quad -9x \\ \hline 3y = -9x - 12 \\ \hline y = -3x - 4 \end{array}$$

$$8) \begin{array}{r} \frac{1}{2}x + y = 9 \\ -\frac{1}{2}x \quad -\frac{1}{2}x \\ \hline y = -\frac{1}{2}x + 9 \end{array}$$

$$4) \begin{array}{r} 7x + 7y = 49 \\ -7x \quad -7x \\ \hline 7y = -7x + 49 \\ \hline y = -x + 7 \end{array}$$

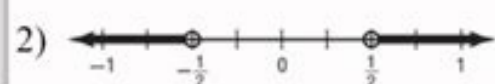
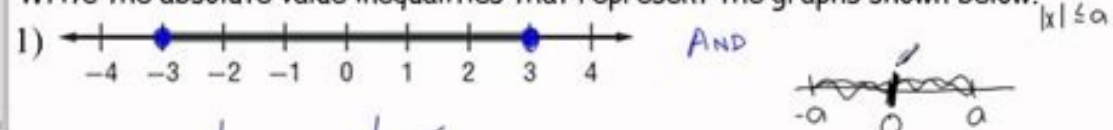
$$9) \begin{array}{r} 2x + 2y + 4 = 0 \\ -2x \quad -2x \\ \hline 2y + 4 = -2x \\ \hline 2y = -2x - 4 \\ \hline y = -x - 2 \end{array}$$

$$5) \begin{array}{r} 4x + y = 10 \\ -4x \quad -4x \\ \hline y = -4x + 10 \end{array}$$

$$10) \begin{array}{r} 2x + 4y = 0 \\ -2x \quad -2x \\ \hline 4y = -2x \\ \hline y = -\frac{1}{2}x \end{array}$$

Exercise Set #6 Writing Absolute Value Inequalities from Graphs

Write the absolute value inequalities that represent the graphs shown below.



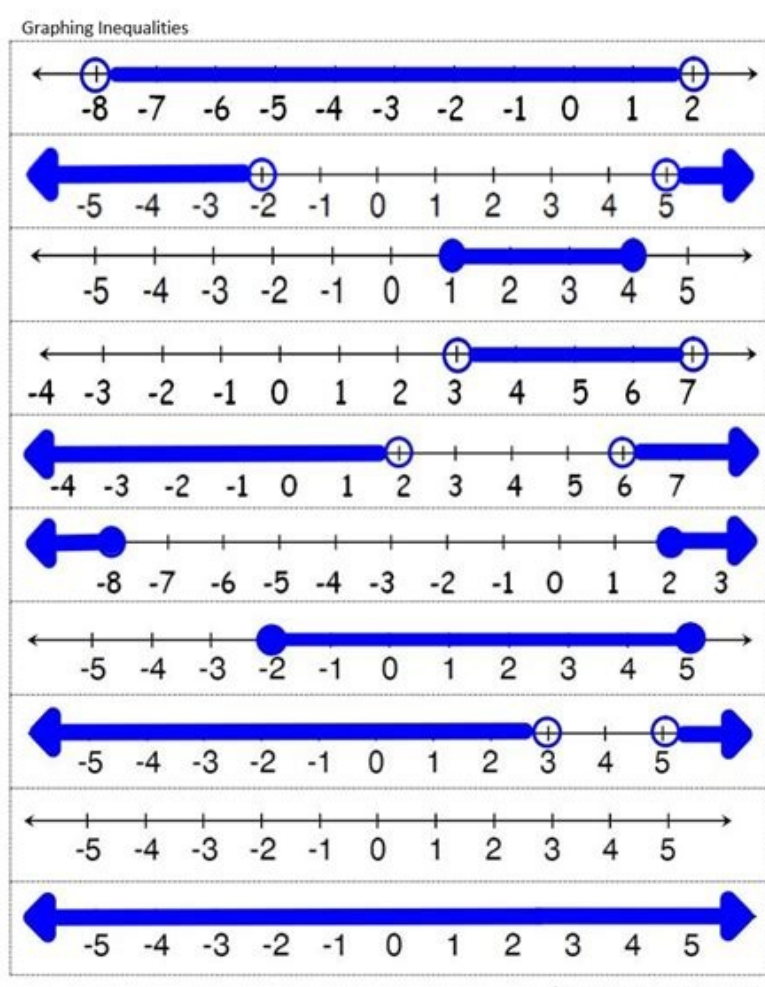
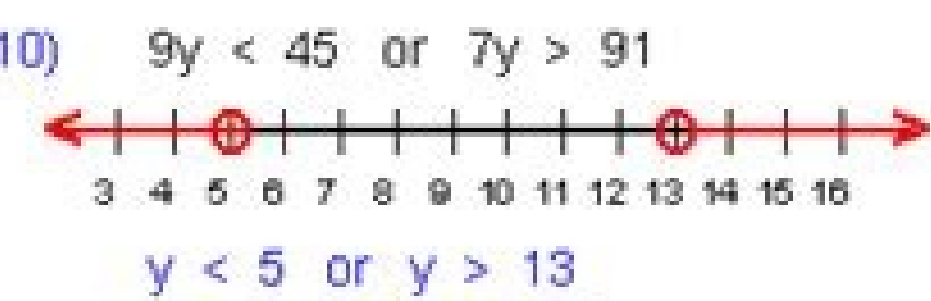
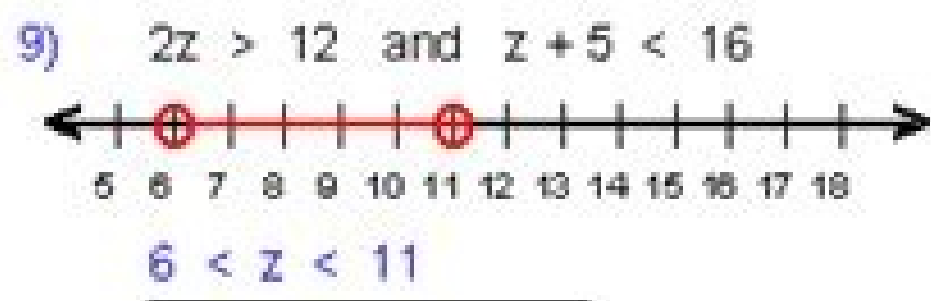
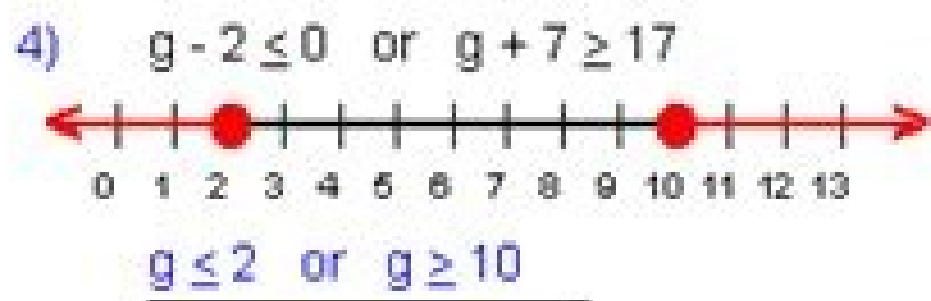
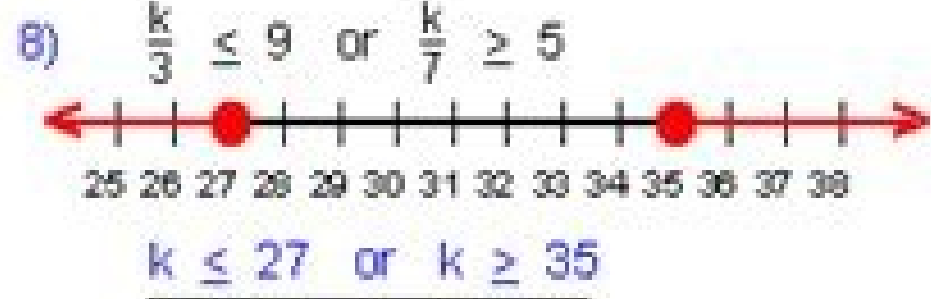
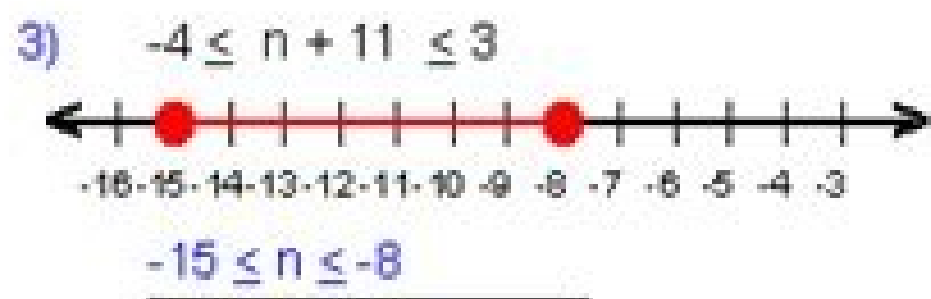
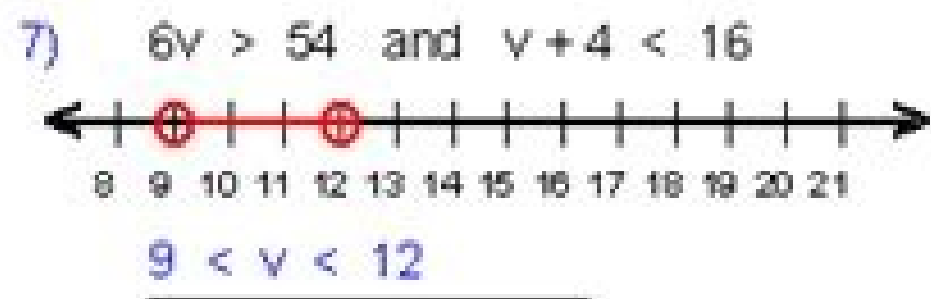
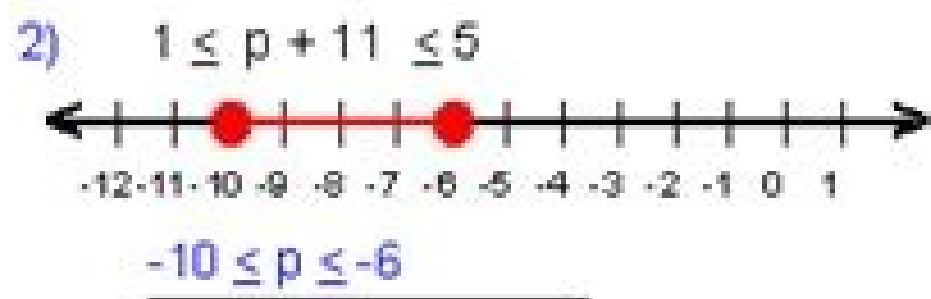
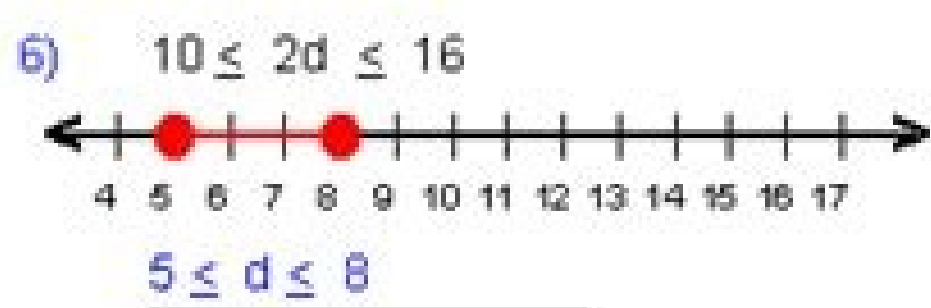
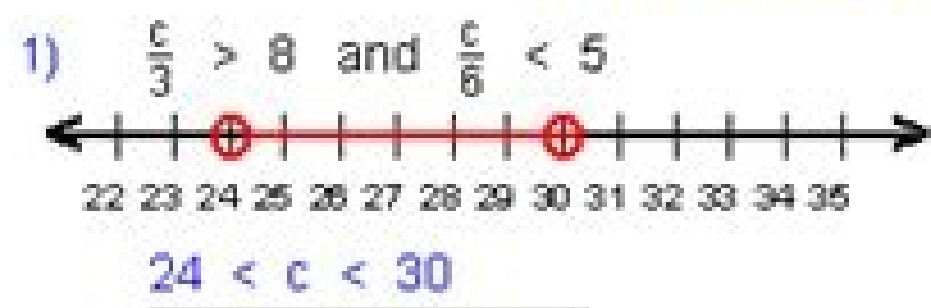
Worksheet: Electron Configurations Name: _____ Period: _____

- Which electron configuration represents an atom in an excited state?
A) $1s^2 2s^2 3p^1$ C) $1s^2 2s^2 2p^2 3p^2$
B) $1s^2 2s^2 2p^2 3p^1$ D) $1s^2 2s^2 2p^2 3s^2$
- All of the elements in Period 3 have a total of 2 electrons in the
A) 2s sublevel C) 2p sublevel
B) 3s sublevel D) 3p sublevel
- Which atom in the ground state has three unpaired electrons in its outermost principal energy level?
A) Li C) N
B) B D) Ne
- What is the total number of valence electrons in an atom with the electron configuration $1s^2 2s^2 2p^3 3s^2 3p^2$?
A) 6 C) 3
B) 2 D) 5
- Which orbital notation represents an atom of beryllium in the ground state?
A) B) C) D)
- Which element has a completely filled third principal energy level?
A) Ar C) Fe
B) N D) Zn
- Which atom in the ground state has only 3 electrons in the 3p sublevel?
A) phosphorus C) argon
B) potassium D) aluminum
- What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?
A) 1 C) 3
B) 2 D) 4
- Which sublevel configuration correctly represents a completely filled third principal energy level?
A) $3s^2 3p^6$ C) $3s^2 3p^5 3d^0$
B) $3s^2 3p^5 3d^0$ D) $3s^2 3p^5 3d^1$
- Which atom in the ground state contains a partially filled 3p orbital?
A) argon C) potassium
B) calcium D) aluminum
- In the ground state, the atoms of elements in Period 2 all have the same number of
A) protons C) 1s electrons
B) neutrons D) oxidation states
- Which is the electron configuration for a neutral atom in the ground state?
A) $1s^2 2s^2 3s^1$ C) $1s^2 2s^2 2p^2 3s^1$
B) $1s^2 2s^2 2p^2 3s^1$ D) $1s^2 2s^2 2p^2 3s^1$
- Which represents the electron configuration of an isotope of oxygen in the ground state?
A) $1s^2 2s^2 2p^2$ C) $1s^2 2s^2 2p^2$
B) $1s^2 2s^2 2p^2$ D) $1s^2 2s^2 2p^2$
- The total number of sublevels in the fourth principal energy level of an atom is
A) 1 C) 3
B) 2 D) 4
- Which is a possible electron configuration for argon in the excited state?
A) $1s^2 2s^2 2p^3 3s^2 3p^1$ C) $1s^2 2s^2 2p^6 3s^1 3p^1$
B) $1s^2 2s^2 2p^3 3s^2 3p^1$ D) $1s^2 2s^2 2p^6 3s^1 3p^1 4s^1$
- Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?
A) All have the same shape but a different orientation.
B) All have the same shape and the same orientation.
C) All have a different shape and a different orientation.
D) All have a different shape but the same orientation.
- A maximum of 6 electrons can occupy
A) an s orbital C) a p orbital
B) an s sublevel D) a p sublevel
- How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?
A) Hydrogen has one electron in a higher energy level.
B) Hydrogen has two electrons in a lower energy level.
C) Hydrogen contains a half-filled orbital.
D) Hydrogen contains a completely filled orbital.

Name : _____ Score : _____

Teacher : _____ Date : _____

Solve and Graph the Inequalities



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Lee Domains of Rational Expressions Worksheet Multiplying Rational Expressions Worksheet Dividing Rational Expressions Notes Divide Rational Expressions Worksheet Dividing Rational Expressions and LCD Worksheet Answers Facing Math Worksheet over Log Functions Properties of Exponents Review GCF and Product Sum Worksheet Key Front of Grouping Worksheet Factor by Grouping Worksheet Key Factoring Quadratics Day 1 Worksheet Key Factoring Quadratics Day 2 Worksheet Key Difference of Two Squares Notes Special Cases Worksheet Key Sum and Difference of Cubes Notes Special Cases #2 Worksheet Key Simplifying Radicals Notes Simplifying Radicals Day 1 Worksheet Key Operations with Radicals Worksheet Answers Rationalizing the Denominator Notes Rationalizing Worksheet Key Rationalizing Higher Order Radicals Notes Rationalizing Higher Order Radicals Worksheet Answers Factoring and Radical Review Complex Numbers Day 1 Worksheet Key Complex Numbers Day 2 Notes Multiplying and Rationalizing Complex Numbers Worksheet Answers Radical and Complex Number Review Interval Notation Notes Key Interval Notation Worksheet Graphing Quadratics - 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